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$P_{\text{vap}} = \exp(14.790 - \frac{5713}{T})$

and by the ideal gas law

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$V = \frac{nRT}{P} = \frac{18 \text{ mol} \times 8.314 \text{ J mol}^{-1} \text{ K}^{-1} \times 10^3 \text{ K}}{1 \times 10^5 \text{ Pa}}$

$\times 10^3 \text{ m}^3 \text{ mol}^{-1}$

$V = 1 \times 10^3 \text{ m}^3$

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? (10 ? x) 10⁶ volume T
taken up by liquid
Equating P and P_{vap}
we have $x \cdot 18 \times 8.314 \times$
 $10^5 \cdot T = 5432.8 \cdot x \cdot I$.

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1.2 An approximate solution can be found if we combine Equations 1.4 and 1.5: $1.2 \text{ mV} \approx k_B T \ln 2$. Assume the temperature is 22 C.

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