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Monatomic Ion. Consists of a single atom with a positive or negative charge resulting from the loss of gain of one or more valence electrons and behaves as a unit, yet it still has a charge. (+ or -). Determined by the amount of

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29 terms. hales913. Chapter 9:
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ion. - single atom with a positive or negative charge resulting from the loss or gain of electrons, respectively.

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Formulas for Acids and Bases.

An acid is a compound that produces H^+ ions when it dissolves in water. The formula for an acid normally starts with an H. When naming acids, you should first determine whether or not the acid contains oxygen.

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Compounds that contain one or more hydrogen atoms and produce.... An ionic compound that produces hydroxide ions when dissolved.... Any atom or group of atoms that has a positive charge. Compounds composed of cations

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Formulas83 SECTION 9.3

NAMING AND WRITING

FORMULAS FOR MOLECULAR

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COMPOUNDS (pages 268 – 270)

This section explains the rules for naming and writing formulas for binary molecular compounds.

Naming Binary Molecular

Compounds (pages 268 – 269) 1.

Circle the letter of the type(s) of elements that form binary

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~~Name Date Class CHEMICAL
NAMES AND FORMULAS 9
SECTION 9.2 NAMING AND
WRITING FORMULAS FOR IONIC
COMPOUNDS I. Write the
formulas for these binary ionic~~

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write the formulas for the

compounds formed c. b. f. 3. Name

the following binary ionic

compounds. sodium sulfide g.

CuCl: h. SnCl_2 . a. MnO_2 c. d. SrBr_2

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Formulas” ... in parenthesis after the name of the metal (Table 9.2, p.255) Predicting Ionic Charges

Some of the post-transition elements also have more than one possible oxidation state. Tin (II) = Sn^{2+} Lead (II) = Pb^{2+} Tin (IV) = Sn^{4+} Lead (IV) = Pb^{4+}

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42. a. 2 –
b. 1+ c. 1 – d. 3+ 43. a. 2+ b. 2+
c. 3+ d. 1+ 44. a. barium ion b.
iodide ion c. silver ion d.
mercury(II) ion 45. cyanide, CN^-
and hydroxide, OH^- 46. a.
hydroxide ion b. lead(IV) ion c.
sulfate ion d. oxide ion 47. zero; A

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Formulas is Electrically Neutral.

48. The symbols for the cation and

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Section 9.1: Naming Ions. 1. What

is the charge on the ion typically

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Answers
formed by each element? a.
oxygen c. sodium e. nickel, two
electrons lost . b. iodine d.

aluminum f. magnesium. 2. How
many electrons does the neutral
atom gain or lose when each ion

forms? a. Cr^{3+} c. Li^{+} e. Cl^{-} b. P^{3-}

...

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negative charge ... place cation
name first followed by the anion
name: naming polyatomic ...

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