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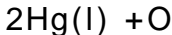
chemical equation

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for the reaction at equilibrium is written using

double arrows to indicate the overall reversibility of the reaction. $2\text{HgO}(s)$



Equilibrium, a Dynamic State

Many chemical reactions are reversible under ordinary conditions

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of temperature and concentration. They will reach a state of equilibrium

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$$10^{-2} \quad K = 3.5 \times 10^4$$

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Section 1

Equilibrium 18.2

Shifting Equilibrium

The Effect of

Change in Pressure:

Ways to Change

Pressure Add or

remove a gaseous

reactant or product

Add an inert gas

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(one not involved in the reaction) An inert gas increases the total pressure but has no effect on the concentrations or partial pressures of the reactants or products

The Effect of Change in Pressure

Change the volume of the container When the volume of the ...

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ANSWER Answer

the following
questions in the
space provided. 1.

Section 1 Raising the
temperature of any
equilibrium system
always (a) favors
the forward
reaction. (b) favors
the reverse
reaction. (c) favors
the exothermic
reaction.

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the exothermic
reaction.

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Section 1

Equilibrium Figure

18-3 If the only way in or out of San Francisco and Sausalito is across the Golden Gate Bridge, which joins the two cities, then the number of

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vehicles in the two cities will remain constant if the number of vehicles per hour crossing the bridge in one direction equals the number of vehi-

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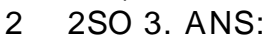
Equilibrium

5. Amongst the following chemical

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reactions, the irreversible reaction is: a)



The reaction is not reversible. 6. When the rate of forward reaction becomes equal to backward

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reaction, this state
is termed as: a)
Chemical
equilibrium. b)
Reversible state

Chemical

Equilibrium

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And Answers

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